**Quaid-e-Azma secondary School Scheme #7 (S.B.B)**

**Class: Pre 9th Subject: Chemistry**

**Chapter 3: Atomic Structure**

**Assessment 1 (Topic 3.1 to 3.4)**

**Total Marks=25**

**Question 1: Choose correct option (1x8=8)**

1.Which particle has no charge?  
A. Proton  
B. Electron  
C. Neutron  
D. Positron

2. Rutherford’s gold foil experiment proved that:  
A. Electrons are positively charged  
B. Atom is mostly empty space  
C. Electrons are present in nucleus  
D. Atoms are indivisible

3. What is the charge on an electron?  
A. +1  
B. 0  
C. -1  
D. +2

4.Atomic number is equal to the number of:  
A. Neutrons  
B. Protons  
C. Electrons + Protons  
D. Nucleons

5.The unit used to express atomic mass is:  
A. Gram  
B. Milligram  
C. Dalton  
D. Atomic mass unit (amu)

6.Which atomic model introduced energy levels or shells?  
A. Dalton’s Model  
B. Rutherford’s Model  
C. Bohr’s Model  
D. Thomson’s Model

7. Relative atomic mass is the average mass of atoms compared to:  
A. Hydrogen atom  
B. Carbon-12 atom  
C. Oxygen atom  
D. Proton

8. What is the approximate mass of a proton in amu?  
A. 0.0005 amu  
B. 0.5 amu  
C. 1 amu  
D. 1.67 amu  
**Question 2: Short Questions (3x3=9)**

1. Define atomic number. How is it related to protons and electrons in a neutral atom?
2. What is charge neutrality?
3. Describe Three postulates of Dalton’s model.

**Question 3: Long Questions (4x2=8)**

1. Write down the conclusions and defects of Rutherford’s model .
2. Atomic number of an element is 17 and mass number is 35. How many protons and neutrons are in the nucleus of an atom of this element?